



5 Factors affecting Rate of Reaction

1. Temperature (any reactions)	↑ temp. → ↑ kinetic energy of particles → ↑ frequency of effective collisions → ↑ rate of rxn
2. Concentration (when one or more reactants are in solutions)	↑ concentration → ↑ number of particles per unit volume → ↑ frequency of effective collisions → ↑ rate of rxn
3. Pressure (when one or more reactants are gases)	↑ pressure of a gas → gas particles are closely packed → ↑ frequency of effective collisions → ↑ rate of rxn
4. Surface Area (when one or more reactants are solids)	↓ particle size of solid → ↑ surface area → ↑ frequency of effective collisions on the surface → ↑ rate of rxn
5. Catalyst	→ Adding a suitable catalyst → Provides an alternative pathway involving a lower activation energy → More colliding particles possess energy equal to or greater than the activation energy → ↑ frequency of effective collisions → ↑ rate of rxn

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